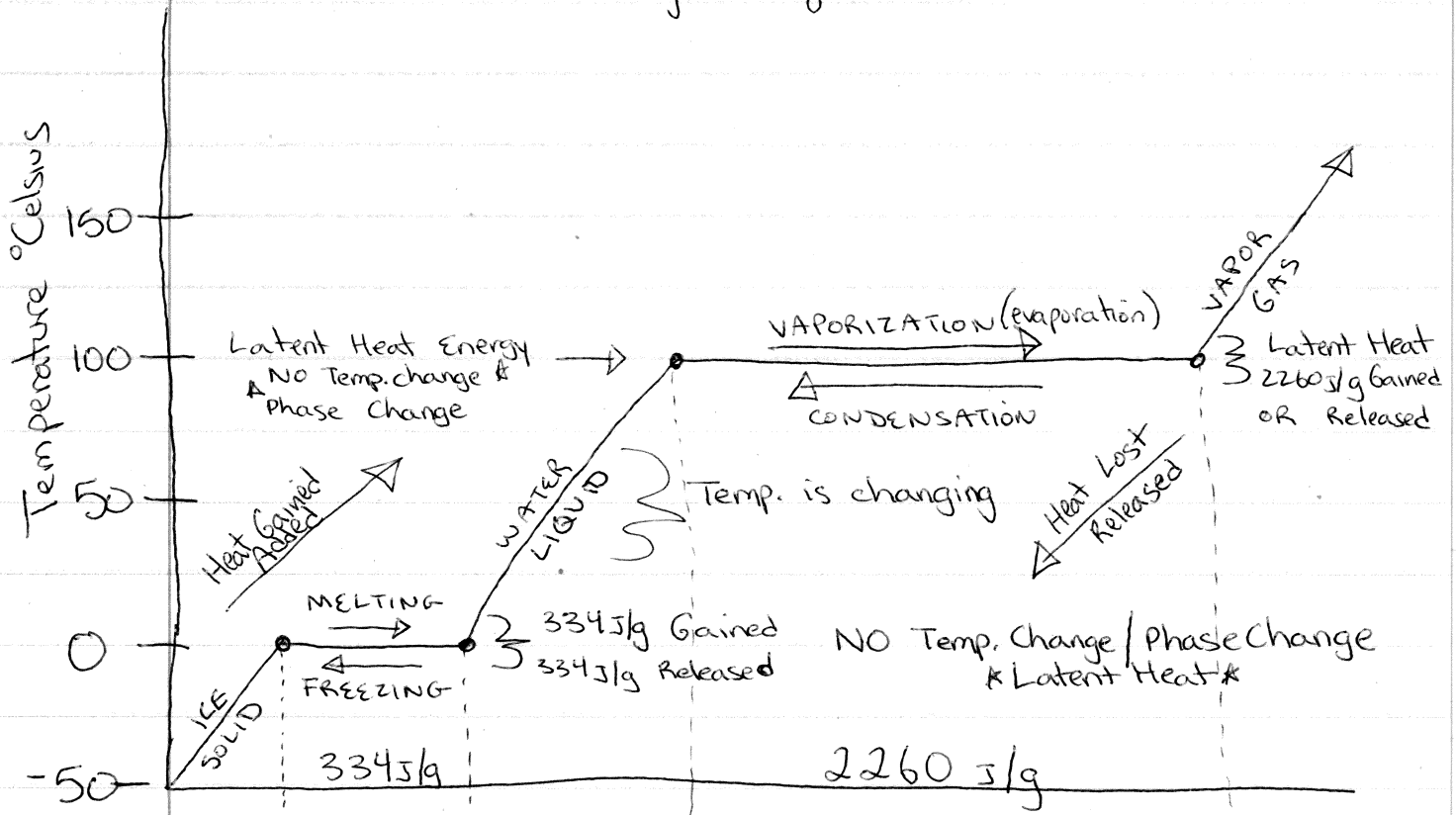


Aim: Describe how water changes phases

	<u>0 minutes</u>	<u>15 minutes</u>	<u>35 minutes</u>
<u>Demo</u>	Ice = -20°C	ice/liquid 0°C melted *	$48^{\circ} = \text{LIQUID now}$
	Liquid Water = 22°C	liquid 60°C	liquid/gas $100^{\circ}\text{C} = \text{Boiling now}$ evaporation beginning

Heating Curve for Water
1 gram liquid water



Heat Energy Added (Joules)

Solid/Ice \rightarrow Liquid = melting 334 J/g Gained \rightarrow

Liquid \rightarrow Gas/Vapor = vaporization 2260 J/g Gained \rightarrow

Gas/Vapor \rightarrow Liquid = condensation 2260 J/g Released \leftarrow Energy enters atmosphere clouds form storms occur

Liquid \rightarrow Solid/Ice = freezing 334 J/g Released \leftarrow LOST to AIR